

name	symbol	relative abundance (on Earth)	# protons	# neutrons	# electrons	comments
Hydrogen	${}^1\text{H}_1$	99.985%	1	0	1	
Deuterium	D (${}^2\text{H}_1$)	0.015%	1	1	1	also called “Heavy Hydrogen”; Deuterium nucleus is known as a “Deuteron”
Tritium	T (${}^3\text{H}_1$)	–	1	2	1	radioactive: half-life=12 yrs.
Helium	${}^3\text{He}_2$	0.0001%	2	1	2	rare!
”	${}^4\text{He}_2$	99.9999%	2	2	2	
Carbon	${}^{12}\text{C}_6$	98.9%	6	6	6	
”	${}^{13}\text{C}_6$	1.1%	6	7	6	
”	${}^{14}\text{C}_6$	–	6	8	6	radioactive: half-life=5700 yrs.
Oxygen	${}^{16}\text{O}_8$	99.76%	8	8	8	
”	${}^{17}\text{O}_8$	0.04%	8	9	8	
”	${}^{18}\text{O}_8$	0.20%	8	10	8	
Iron	${}^{54}\text{Fe}_{26}$	5.8%	26	28	26	
”	${}^{56}\text{Fe}_{26}$	91.8%	26	30	26	
”	${}^{57}\text{Fe}_{26}$	2.1%	26	31	26	
”	${}^{58}\text{Fe}_{26}$	0.3%	26	32	26	
”	${}^{60}\text{Fe}_{26}$	–	26	34	26	radioactive: half-life=1.5 million yrs.
Uranium	${}^{235}\text{U}_{92}$	0.7%	92	143	92	radioactive: half-life=700 million yrs.; can also fission
”	${}^{238}\text{U}_{92}$	99.3%	92	146	92	radioactive: half-life=4.5 billion yrs.

Table 1: Some elements and their isotopes. The “half-life” of a radioactive isotope is the time it takes on average for half the atoms in a sample to decay.